



MARCH 22, 2026

CHEMISTRY
MCQ PRACTICE SERIES

Daily Questions for Competitive Exam - 16 Jan 2026 to 31 Jan 2026 MCQs

GoPract



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NEET Chemistry MCQs – Polymers - 31 Jan 2026

1. Which of the following is a natural polymer?

- A) Nylon-6
- B) Bakelite
- C) Cellulose
- D) PVC

2. Which polymer is formed by addition polymerisation?

- A) Nylon-6,6
- B) Terylene
- C) Polythene
- D) Bakelite

3. Which polymer is used to make non-stick cookware?

- A) PVC
- B) Teflon
- C) Polystyrene
- D) Nylon

4. Which of the following is a biodegradable polymer?

- A) Polythene
- B) PVC
- C) Nylon
- D) PHBV

5. Bakelite is an example of:

- A) Elastomer
- B) Thermoplastic
- C) Thermosetting polymer
- D) Fibre

Answer Key

1. C
2. C
3. B
4. D
5. C

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Expert Tip – Polymers (Easy Memory)

Remember this simple NEET rule: **Addition polymer** → **one monomer**. **Condensation polymer** → **two monomers + water loss**. Natural polymers come from plants or animals. Bakelite never melts again because it is thermosetting. Keep these points in mind and you can solve most **chem mcq** from Polymers easily.

NEET Chemistry MCQs – Amines - 30 Jan 2026

1. Which amine is the strongest base in aqueous solution?
 - A) Methylamine
 - B) Dimethylamine
 - C) Trimethylamine
 - D) Aniline
2. Which test is used to distinguish primary, secondary and tertiary amines?
 - A) Tollen's test
 - B) Fehling's test
 - C) Hinsberg test
 - D) Iodoform test
3. Which amine gives a positive carbylamine test?
 - A) Primary amine
 - B) Secondary amine
 - C) Tertiary amine
 - D) Aromatic nitro compound
4. Aniline is less basic than methylamine because:
 - A) Aniline has larger size
 - B) Lone pair is involved in resonance
 - C) Aniline is aromatic
 - D) Methylamine has no lone pair
5. Which compound is formed when aniline reacts with nitrous acid at 0–5°C?
 - A) Nitrobenzene
 - B) Phenol
 - C) Benzene diazonium chloride
 - D) Chlorobenzene

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Answer Key

1. B
2. C
3. A
4. B
5. C

Expert Tip – Amines (Easy Memory Rule)

Remember this simple NEET rule: **2° amine > 1° amine > 3° amine** in basic strength in water. Only **primary amines** give carbylamine test. Aniline is weak because its lone pair joins the benzene ring. If you remember these three points, you can solve most **chem mcq** from Amines without confusion.

NEET Chemistry MCQs – Aldehydes, Ketones and Carboxylic Acids - 29 Jan 2026

1. Which compound gives a positive Tollens' test?

- A) Acetone
- B) Benzophenone
- C) Acetaldehyde
- D) Acetic acid

2. Which functional group is present in carboxylic acids?

- A) –CHO
- B) –CO–
- C) –COOH
- D) –OH

3. Which compound shows the highest boiling point?

- A) Methanal
- B) Ethanol
- C) Ethanoic acid
- D) Propanone

4. Which reagent converts aldehydes into carboxylic acids?

- A) NaBH₄
- B) PCC
- C) Tollens' reagent
- D) KMnO₄

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5. Which compound does NOT show iodoform test?

- A) Ethanol
- B) Acetone
- C) Acetaldehyde
- D) Methanol

Answer Key

1. C
2. C
3. C
4. D
5. D

Expert Tip – Aldehydes, Ketones & Carboxylic Acids

Remember this NEET rule: **Aldehyde = Oxidation easy, Ketone = Oxidation hard.** Carboxylic acids have the highest boiling point due to strong hydrogen bonding. If you remember Tollens', Fehling's, and iodoform tests clearly, you can solve most **chem mcq** from this chapter without confusion.

NEET Chemistry MCQs – Alcohols, Phenols and Ethers - 28 Jan 2026

1. Which alcohol shows the highest boiling point?

- A) Methanol
- B) Ethanol
- C) Propan-1-ol
- D) Propan-2-ol

2. Phenol is more acidic than ethanol because phenoxide ion is:

- A) Less stable
- B) Stabilized by resonance
- C) Smaller in size
- D) Neutral in nature

3. Which reagent converts phenol into tribromophenol?

- A) Bromine water
- B) Br₂ in CCl₄
- C) HBr
- D) PBr₃

4. Which ether forms explosive peroxides on standing in air?

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- A) Diethyl ether
- B) Dimethyl ether
- C) Anisole
- D) Phenyl ether

5. Which reaction is used to prepare ethers from alcohols?

- A) Friedel–Crafts reaction
- B) Williamson synthesis
- C) Kolbe reaction
- D) Reimer–Tiemann reaction

Answer Key

1. C
2. B
3. A
4. A
5. B

Expert Note – Alcohols, Phenols and Ethers (NEET)

NEET often tests acidity order, boiling point trends, and named reactions from this chapter. Phenol is more acidic due to resonance stabilization. Primary alcohols show higher boiling points due to stronger hydrogen bonding. Williamson synthesis is the most important method to form ethers. Practice such **chem mcq** regularly to answer chemistry questions with confidence in the exam.

NEET Chemistry MCQs – Haloalkanes and Haloarenes - 27 Jan 2026

1. Which bond breaks during a nucleophilic substitution reaction in haloalkanes?

- A) C–C bond
- B) C–H bond
- C) C–X bond
- D) X–X bond

2. Which alkyl halide undergoes S_N1 reaction most easily?

- A) CH_3Cl
- B) C_2H_5Cl
- C) $(CH_3)_3CCl$
- D) $CH_3CH_2CH_2Cl$

3. Which solvent favors the S_N1 mechanism?

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- A) Hexane
- B) Ether
- C) Water
- D) Benzene

4. Which halogen atom forms the strongest bond with carbon?

- A) F
- B) Cl
- C) Br
- D) I

5. Which compound is an example of a haloarene?

- A) CH₃Cl
- B) C₂H₅Br
- C) Chlorobenzene
- D) CH₂Cl₂

Answer Key

1. C
2. C
3. C
4. A
5. C

Expert Note – Haloalkanes and Haloarenes

In NEET **chem mcq**, this chapter tests reaction mechanism. Remember the order: tertiary > secondary > primary for S_N1. Polar solvents support S_N1 reactions. Strong C–F bond makes fluorides less reactive. Practice such **chemistry questions and answers** daily on GoPract.

NEET Chemistry MCQs – Hydrocarbons - 26 Jan 2026

1. Which hydrocarbon shows aromatic character?
 - A) Ethene
 - B) Ethyne
 - C) Benzene
 - D) Cyclohexane
2. Which reagent is used to distinguish between alkanes and alkenes?
 - A) FeCl_3
 - B) Bromine water
 - C) Tollens reagent
 - D) Sodium metal
3. The general formula of alkynes is:
 - A) $\text{C}_n\text{H}_{2n+2}$
 - B) C_nH_{2n}
 - C) $\text{C}_n\text{H}_{2n-2}$
 - D) C_nH_n
4. Which reaction is shown by alkanes mainly?
 - A) Addition reaction
 - B) Elimination reaction
 - C) Substitution reaction
 - D) Polymerisation
5. Which hydrocarbon gives a sooty flame on burning?
 - A) Methane
 - B) Ethane
 - C) Ethene
 - D) Benzene

Answer Key

1. C
2. B
3. C
4. C
5. D

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Expert Note – Hydrocarbons (NEET)

In NEET **chem mcq**, hydrocarbons test basic concepts. Focus on general formula, reactions, and flame test. Aromatic compounds give sooty flame. Practice such **chemistry questions and answers** daily. This habit helps you answer chemistry questions with confidence in exam.

NEET Chemistry MCQs – General Organic Chemistry (GOC) - 25 Jan 2026

1. Which effect explains the permanent displacement of electrons in a covalent bond?

- A) Inductive effect
- B) Resonance effect
- C) Hyperconjugation
- D) Electromeric effect

2. Which carbocation is the most stable?

- A) CH_3^+
- B) C_2H_5^+
- C) $(\text{CH}_3)_2\text{CH}^+$
- D) $(\text{CH}_3)_3\text{C}^+$

3. Which of the following shows a +I effect?

- A) $-\text{NO}_2$
- B) $-\text{CN}$
- C) $-\text{CH}_3$
- D) $-\text{COOH}$

4. Which species acts as a nucleophile?

- A) BF_3
- B) AlCl_3
- C) OH^-
- D) H^+

5. Which type of reaction involves movement of a pair of electrons?

- A) Free radical reaction
- B) Electrophilic reaction
- C) Nucleophilic reaction
- D) Ionic reaction

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Answer Key

1. B
2. D
3. C
4. C
5. D

Expert Tip – General Organic Chemistry

GOC builds the base of organic chemistry. NEET asks logic, not memory. Focus on effects like inductive and resonance. Compare stability step by step. Daily MCQ practice on GoPract helps in exam preparation and improves your speed when you study for exam.

NEET Chemistry MCQs – Environmental Chemistry - 24 Jan 2026

1. Which gas is the main cause of acid rain?

- A) Carbon monoxide
- B) Sulphur dioxide
- C) Methane
- D) Nitrogen

2. Which pollutant causes photochemical smog?

- A) Ozone
- B) Sulphur dioxide
- C) Carbon dioxide
- D) Lead

3. Which layer of the atmosphere absorbs harmful UV rays?

- A) Troposphere
- B) Mesosphere
- C) Stratosphere
- D) Thermosphere

4. Which chemical is responsible for ozone layer depletion?

- A) Carbon dioxide
- B) Chlorofluorocarbons
- C) Sulphur dioxide
- D) Nitrous oxide

5. Which gas is known as a greenhouse gas?

- A) Oxygen
- B) Nitrogen

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- C) Carbon dioxide
- D) Argon

Answer Key

1. B
2. A
3. C
4. B
5. C

Expert Tip – Environmental Chemistry

In NEET **Exam Preparation**, this chapter gives direct questions. Read NCERT lines carefully. Focus on gases, layers, and causes. Do not guess. Revise facts daily. Practice such MCQs on **GoPract** to **study for exam** with confidence.

NEET Chemistry MCQs – Some p-Block Elements - 23 Jan 2026

1. Which group 15 element shows maximum tendency to form multiple bonds?

- A) Nitrogen
- B) Phosphorus
- C) Arsenic
- D) Bismuth

2. Which oxide of nitrogen is neutral in nature?

- A) N_2O
- B) NO_2
- C) N_2O_5
- D) NO

3. Which allotrope of phosphorus is the most reactive?

- A) White phosphorus
- B) Red phosphorus
- C) Black phosphorus
- D) Violet phosphorus

4. Which halogen has the highest electronegativity?

- A) Chlorine
- B) Bromine
- C) Fluorine
- D) Iodine

5. Which acid is formed when nitrogen dioxide reacts with water?

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- A) Nitric acid only
- B) Nitrous acid only
- C) Both nitric and nitrous acid
- D) No acid is formed

Answer Key

1. A
2. D
3. A
4. C
5. C

Expert Tip – p-Block Elements

In NEET Exam Preparation, p-Block questions often come directly from NCERT facts. Focus on trends, oxides, and allotropes. Revise reactions daily. This habit helps you **study for exam** with confidence. Practice MCQs on GoPract to avoid silly mistakes.

NEET Chemistry MCQs – The s-Block Elements - 22 Jan 2026

1. Which of the following alkali metals has the **highest hydration enthalpy**?

- A) Li
- B) Na
- C) K
- D) Cs

2. Which alkaline earth metal sulphate is **insoluble in water**?

- A) MgSO_4
- B) CaSO_4
- C) SrSO_4
- D) BaSO_4

3. The anomalous behaviour of lithium compared to other alkali metals is mainly due to:

- A) High atomic mass
- B) Low ionisation energy
- C) Small size and high polarising power
- D) Presence of d-orbitals

4. Which of the following compounds is used in the **Solvay process** for the manufacture of sodium carbonate?

- A) NaCl
- B) NaOH
- C) NaHCO_3

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- D) Na_2SO_4

5. Which of the following statements about alkaline earth metals is correct?

- A) They form monovalent ions
- B) Their oxides are acidic
- C) They are less reactive than alkali metals
- D) They do not form complexes

Answer Key

1. A
2. D
3. C
4. C
5. C

Expert Tip – s-Block Elements (NEET)

In NEET **Exam Preparation**, s-Block questions are often **NCERT-line based**. Focus on trends like hydration enthalpy, solubility of sulphates and anomalous behaviour of lithium. Regular MCQ practice on GoPract helps you **study for exam** smartly and avoid common trap options.

NEET Chemistry MCQs – Hydrogen (Medium Level) - 21 Jan 2026

1. Which of the following hydrides is **saline in nature**?

- A) CH_4
- B) NH_3
- C) NaH
- D) B_2H_6

2. Which of the following metals reacts with hydrogen to form an **interstitial hydride**?

- A) Sodium
- B) Magnesium
- C) Palladium
- D) Calcium

3. Heavy water (D_2O) is mainly used:

- A) As a reducing agent
- B) As a moderator in nuclear reactors
- C) For extraction of metals
- D) As a solvent in organic reactions

4. Which of the following statements regarding hydrogen peroxide is correct?

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- A) It acts only as an oxidising agent
- B) It acts only as a reducing agent
- C) It can act as both oxidising and reducing agent
- D) It is neutral in all reactions

5. The oxidation state of oxygen in hydrogen peroxide (H_2O_2) is:

- A) -2
- B) -1
- C) 0
- D) +1

Answer Key

1. C
2. C
3. B
4. C
5. B

Short Note – Hydrogen (NEET)

Hydrogen forms three main types of hydrides: **ionic (saline)**, **covalent** and **interstitial**. Hydrogen peroxide is a special compound where oxygen has an oxidation state of **-1**, allowing it to behave as both an oxidising and a reducing agent — a frequently tested NEET concept.

NEET Chemistry MCQs – Redox Reactions (High Difficulty) - 20 Jan 2026

1. In the reaction $\text{MnO}_4^- + 8\text{H}^+ + 5\text{e}^- \rightarrow \text{Mn}^{2+} + 4\text{H}_2\text{O}$, the number of electrons gained per manganese atom is:

- A) 3
- B) 4
- C) 5
- D) 7

2. Which of the following species acts as an **oxidising agent** in the reaction: $\text{Zn} + \text{Cu}^{2+} \rightarrow \text{Zn}^{2+} + \text{Cu}$?

- A) Zn
- B) Zn^{2+}
- C) Cu
- D) Cu^{2+}

3. The oxidation number of sulphur in sodium thiosulphate ($\text{Na}_2\text{S}_2\text{O}_3$) is:

- A) +2

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- B) +5
- C) -2
- D) Average value of +2

4. Which of the following reactions is an example of a **disproportionation reaction**?

- A) $\text{Zn} + \text{CuSO}_4 \rightarrow \text{ZnSO}_4 + \text{Cu}$
- B) $2\text{H}_2\text{O} \rightarrow 2\text{H}_2 + \text{O}_2$
- C) $\text{Cl}_2 + 2\text{OH}^- \rightarrow \text{Cl}^- + \text{ClO}^- + \text{H}_2\text{O}$
- D) $\text{Fe} + \text{S} \rightarrow \text{FeS}$

5. In acidic medium, which of the following is the correct balancing factor for electrons while balancing the reaction between Fe^{2+} and dichromate ion ($\text{Cr}_2\text{O}_7^{2-}$)?

- A) 3
- B) 5
- C) 6
- D) 7

Answer Key

1. C
2. D
3. D
4. C
5. C

Short Note – Redox Reactions (NEET)

Oxidation number is a powerful tool to identify oxidation and reduction. In complex ions like thiosulphate, individual atoms may have different oxidation states, so an **average value** is used. Disproportionation reactions involve the same element undergoing both oxidation and reduction — a very common NEET trap.

NEET Chemistry MCQs – Equilibrium - 19 Jan 2026

1. For the reaction $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$, if the equilibrium constant K_c is very large, it indicates that:

- A) Reactants are favoured
- B) Products are favoured
- C) Reaction does not proceed
- D) Reaction is very slow

2. Which of the following changes will **increase the yield of products** for an exothermic reaction at equilibrium?

- A) Increase in temperature
- B) Decrease in temperature
- C) Addition of catalyst
- D) Decrease in pressure

3. The unit of equilibrium constant K_c for the reaction $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$ is:

- A) mol L^{-1}
- B) $\text{mol}^{-1} \text{L}$
- C) $\text{mol}^{-2} \text{L}^2$
- D) No unit

4. Which of the following statements is correct for a catalyst in an equilibrium reaction?

- A) It changes the value of K_c
- B) It shifts equilibrium towards products
- C) It increases the rate of both forward and backward reactions equally
- D) It decreases activation energy of forward reaction only

5. For the equilibrium $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$, if the initial concentrations of H_2 and I_2 are equal, then at equilibrium:

- A) $[\text{H}_2] = [\text{I}_2]$
- B) $[\text{HI}] = [\text{H}_2]$
- C) $[\text{HI}] = [\text{I}_2]$
- D) $[\text{H}_2] \neq [\text{I}_2]$

Answer Key

1. B
2. B
3. B
4. C
5. A

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Short Note – Equilibrium (NEET)

The equilibrium constant (K_c) depends only on temperature. A catalyst does not change the position of equilibrium or the value of K_c ; it only helps the system reach equilibrium faster by increasing the rate of both forward and backward reactions.

NEET Chemistry MCQs – Thermodynamics - 18 Jan 2026

1. Which of the following quantities is a **state function**?

- A) Heat
- B) Work
- C) Internal energy
- D) Path length

2. For an **endothermic reaction** carried out at constant pressure, which of the following is correct?

- A) $\Delta H < 0$
- B) $\Delta H > 0$
- C) $\Delta G = 0$
- D) $\Delta S < 0$

3. The relation between enthalpy change (ΔH) and internal energy change (ΔU) for a reaction involving change in gaseous moles is:

- A) $\Delta H = \Delta U - \Delta nRT$
- B) $\Delta H = \Delta U + \Delta nRT$
- C) $\Delta H = \Delta U \times \Delta nRT$
- D) $\Delta H = \Delta U / \Delta nRT$

4. Which of the following processes is **spontaneous** at all temperatures?

- A) $\Delta H > 0, \Delta S < 0$
- B) $\Delta H < 0, \Delta S > 0$
- C) $\Delta H > 0, \Delta S > 0$
- D) $\Delta H < 0, \Delta S < 0$

5. The efficiency of a reversible heat engine operating between two temperatures depends on:

- A) Nature of working substance
- B) Amount of heat absorbed
- C) Temperatures of source and sink
- D) Time taken for the cycle

Answer Key

1. C

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2. B
3. B
4. B
5. C

NEET Exam Tip – Thermodynamics

Always remember: $\Delta G = \Delta H - T\Delta S$. For spontaneity at all temperatures, the condition must be $\Delta H < 0$ and $\Delta S > 0$ — this concept is repeatedly tested in NEET through tricky options.

NEET Chemistry MCQs – States of Matter (Gases & Liquids) - 17 Jan 2026

1. Which of the following gases deviates most from ideal gas behaviour at high pressure and low temperature?

- A) H_2
- B) N_2
- C) CO_2
- D) He

2. The value of compressibility factor (Z) for an ideal gas is:

- A) 0
- B) Less than 1
- C) Greater than 1
- D) Equal to 1

3. Which of the following factors is responsible for the high boiling point of water?

- A) High molecular mass
- B) Hydrogen bonding
- C) Dipole moment
- D) van der Waals forces

4. At constant temperature, the pressure of a gas is doubled. What happens to the volume of the gas?

- A) Volume becomes double
- B) Volume becomes half
- C) Volume remains constant
- D) Volume becomes four times

5. Which of the following statements is correct regarding real gases?

- A) They obey Boyle's law at all temperatures
- B) They show no intermolecular attraction
- C) They liquefy at very low temperatures
- D) Their molecular volume is negligible

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Answer Key

1. C
2. D
3. B
4. B
5. C

NEET Exam Tip – States of Matter

Real gases deviate most from ideal behaviour at **high pressure and low temperature**. Remember: gases with **strong intermolecular forces** (like CO_2) deviate more and are easier to liquefy — a very common NEET trap.

NEET Chemistry MCQs – Chemical Bonding & Molecular Structure - 16 Jan 2026

Chemical Bonding Medium

1. Which of the following molecules has **zero dipole moment**?

- A) NH_3
- B) H_2O
- C) CO_2
- D) SO_2

VSEPR Theory Medium

2. The molecular shape of ammonia (NH_3) is:

- A) Tetrahedral
- B) Trigonal planar
- C) Trigonal pyramidal
- D) Linear

Hybridisation Medium

3. The central atom in SF_6 undergoes which type of hybridisation?

- A) sp^3
- B) sp^3d
- C) sp^3d^2
- D) sp^2d

Hydrogen Bonding Medium

4. Which of the following substances shows the **strongest hydrogen bonding**?

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- A) H₂O
- B) NH₃
- C) HF
- D) CH₃OH

Molecular Orbital Theory Medium

5. The bond order of the oxygen molecule (O₂) according to molecular orbital theory is:

- A) 1
- B) 1.5
- C) 2
- D) 2.5

Answer Key

1. C
2. C
3. C
4. C
5. C

Key Notes – Chemical Bonding & Molecular Structure

- Symmetrical linear molecules like CO₂ have zero dipole moment.
- NH₃ has trigonal pyramidal shape due to one lone pair.
- SF₆ shows sp³d² hybridisation and octahedral geometry.
- Hydrogen bonding strength is maximum in HF due to high electronegativity of fluorine.
- Bond order = (bonding electrons – antibonding electrons) / 2.